

Chemistry 12 - Unit 3 KEY Solubility

14. You are given a bottle of saturated calcium chloride solution at 25°C. Outline the experimental procedure you would use to determine the solubility of CaCl_2 in grams. (3 marks)

3

- weigh empty dish
- add (10.0 mL) of sat. CaCl_2 soln.
- evaporate
- weigh solid CaCl_2 (s)
- Calculate sol. in g/L

15. A solution of FeBr_3 has a $[\text{Br}^-]$ which is 0.36 M. Find the $[\text{Fe}^{3+}]$ in the same solution. (1 mark)

1

Answer: 0.12 M

16. You dissolve 3.28×10^3 moles of $\text{Fe}(\text{NO}_3)_3$ into water to make 2.00 L of solution.

2

17. Determine the $[\text{Fe}^{3+}]$ (1 mark) Answer: 0.0640 M ($1.6 \times 10^{-2} \text{ M}$)

18. Determine the $[\text{NO}_3^-]$ (1 mark) Answer: 0.0960 M ($2.4 \times 10^{-2} \text{ M}$)

19. 30.0 mL of 0.850 M $\text{Ba}(\text{NO}_3)_2$ is added to 50.0 mL of water. The volume is 80.0 mL in the mixture.

2

20. Find the new $[\text{Ba}(\text{NO}_3)_2]$ in the solution. (2 marks)

$0.850 \text{ M} \times \frac{30.0 \text{ mL}}{80.0 \text{ mL}} = 0.31875 \text{ M}$

Answer: 0.319 M

21. Find the new $[\text{NO}_3^-]$ in the solution. (1 mark)

1

Answer: 0.638 M

22. 67.0 mL of 0.25 M BaCl_2 is mixed with 25.0 mL of 0.50 M NaCl . Determine the final $[\text{Cl}^-]$ in the new mixture. (2 marks)

2

$[\text{BaCl}_2] = 0.25 \text{ M} \times \frac{67.0 \text{ mL}}{92.0 \text{ mL}} = 0.18261 \text{ M}$ ($[\text{Cl}^-] = 0.36522 \text{ M}$)

$[\text{NaCl}] = 0.50 \text{ M} \times \frac{25.0 \text{ mL}}{92.0 \text{ mL}} = 0.09891 \text{ M}$ Answer: 0.46413 M

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